

EXPERIMENT 1

ACID BASE TITRATION

Name
Section / Course
I.C. No

Objective: :

.....

.....

Data:

Mass of empty weighing bottle :

Mass of empty weighing bottle and

Potassium hydrogen phthalate

$\text{HKC}_8\text{H}_4\text{O}_4$:

Mass of $\text{HKC}_8\text{H}_4\text{O}_4$ crystals :

Volume of solution prepared :

Questions:

Part A .

1. Show all work in calculating the concentration of the potassium hydrogen phthalate, $\text{HKC}_8\text{H}_4\text{O}_4$, solution in mol/L.

Part B.

Titration data:

Volume of acid pipetted: _____

Buret readings	1	2	3
Final reading (mL)			
Initial reading (mL)			
Volume of NaOH			
Average buret reading	$\frac{\text{Reading 2} + 3}{2}$		

2. Write the balanced equation for the titration.

3. Show all work in calculating the concentration of the NaOH solution in mol/L.

4. Explain why the average buret reading is taken from the last 2 trials?