

EXPERIMENT 2

STOICHIOMETRY : PRECIPITATION OF LEAD SULFATE

Name
Section / Course
Partner's name (if applicable)	

Objective :

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Data :

Volume of solution $\text{Pb}(\text{NO}_3)_2$ = mL

Weight of crucible = g

Weight of crucible + PbSO_4 (weight 1) = g

Weight of crucible + PbSO_4 (weight 2) = g

Weight of crucible + PbSO_4 (weight 3) = g

Weight of precipitate PbSO_4 = g

Observation

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Questions

1. Write a balance chemical reaction and write the mol and mass relationship in the experiment.
2. Calculate the mass of $\text{Pb}(\text{NO}_3)_2$ in the reaction.
3. Calculate the number of mol of $\text{Pb}(\text{NO}_3)_2$ that have reacted.
4. Calculate the molarity of $\text{Pb}(\text{NO}_3)_2$ solution.

5. Name the limiting reagent in the reaction. State your reason.

6. Explain the experimental error that might affect the experimental result.